# **PURITY AND HEAT OF FUSION DATA FOR ENVIRONMENTAL STANDARDS AS DETERMINED BY DIFFERENTIAL SCANNING CALORIMETRY**

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#### ABSTRACT

Differential scanning calorimetry (DSC) has been applied to 273 environmental standards, including pesticides, herbicides and related compounds. Members of the following chemical classes were analyzed: organophosphorus, organochlorine, phenol, triazine, uracil, phenoxy acid, urea, carboxylic acid, amide, and others including amines, organometallics, esters and heterocycles. Values for the heat of fusion, experimental temperature onset, theoretical temperature onset for 100% pure compound, and percent purity are presented. DSC was found to be a widely applicable method to most classes of organic environmental standards and their metabolites.

#### INTRODUCTION

**DSC has been in use for almost 30 years in the routine analysis of high-purity compounds such as pharmaceuticals and for quality control in the manufacture of plastics [l-4]. The theory and practice of DSC have been reviewed and summarized [5,6]. Calorimetric data for certain pesticides are available [7,8]. Heat of fusion values for environmental standards have generated limited interest in the past and, in general, the manufacturers of commercial chemicals do not necessarily require the stringent control of purity ( > 97%) needed for the application of DSC.** 

**The U.S. Environmental Protection Agency (EPA) maintains a repository of pesticides and industrial chemicals to provide high-purity standards and** 

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reference materials to the scientific community. A number of these reference materials are chemicals regulated under the U.S. EPA Resource Conservation and Recovery Act (RCRA). Many of these compounds and their metabolites are available at sufficiently high purity levels to permit analysis by the DSC technique. The wide variety of complex organic structures can cause non-routine thermal behavior (e.g. decomposition, formation of solid solutions, solid-solid phase transitions, and sublimation) difficult to predict but analytically significant [7,8]. In addition, the use of  $\Delta H_f$  has been applied as an essential parameter for relating the solubility of solids in a supercritical fluid [9]. DSC provides an accurate measurement of  $\Delta H$ , through analysis of the melting curve. Compound purity is determined quantitatively by DSC, and identity is checked qualitatively by melting point.

Even when the purity cannot be determined, it is occasionally possible to gain qualitative evidence by DSC to assist in the identification of the compound. As examples, some compounds decompose rather than melt, and others such as strychnine exhibit a characteristic and unusual melting range. The purpose of this paper is to report the heat of fusion  $(\Delta H_f)$ , melting point, and purity data determined by DSC for 273 environmental standards in support of the EPA Quality Assurance and Quality Control Program. The data are valuable to the EPA both as an analytical reference point and as a source for physical and thermodynamic properties.

#### EXPERIMENTAL

The DSC system consisted of a Perkin-Elmer DSC-2 coupled to a Perkin-Elmer model 3600 data station using Thermal Analysis Data Station (TADS) software. The instrument was calibrated daily with a high-purity indium standard (99.9999%). The acceptance limits for heat of fusion and temperature onset values were set at  $6.8 \pm 0.2$  cal g<sup>-1</sup> and 429.79 K, respectively. These control limits were set because of the slow drift in temperature onset over time for this instrument.

All samples were obtained from the Pesticides and Industrial Chemicals Repository maintained by the EPA. Samples were prepared by grinding to a fine powder and weighing 1.5 to 2.0 mg  $\pm$  0.2 mg into an aluminum pan on a Cahn electrobalance. An aluminum cap was placed on the pan and the cap and pan were coldpressed together to form a pellet. Note that gold pans and caps were used for compounds which could react with aluminum, such as organomercury standards. A minimum of two pellets were run for each pesticide. In general, samples were heated at a rate of 2.5 K min<sup>-1</sup>, over a 20 K range, from an initial temperature of 10 K below the expected temperature onset. The sample weight was selected to produce the best accuracy by providing optimum thermal contact and heat transfer within the sample pan.

#### **TABLE 1**

#### DSC results on environmental standards



















 $\sim 10^{-1}$ 



Oxyfluorfen 42874-03-3 2-chloro-1-(3-ethoxy-4-nitrophenoxy)-4-(trifluoromethyl)benzene<br>99.39 7186 358.78 359.75 357 [15] 99.39 7186 358.78 359.75 357 [15] Pentachloraniline 527-20-8<br>99.66 4469 505.78 507.06 99.66 4469 505.78 507.06 505 [18]<br>midone 32809-16-8 Procymidone 3-(3,5-dichloropheny1)-1,5-dimethyl-3-azabicyclo[3.1.0] hexane-2,4-dione 99.78 7191 438.15 438.40 439 [12] Prolan 117-27-1 l,l'-(2nitropropylidene)bis(4-chlorobenzene) 97.9 5112 354.26 355.17 354 [15] Pyrazon (chloridazon) 1698-60-8 5-amino-4-chloro-2-phenyl-3(2 $H$ )-pyridazinone 99.90 6394 479.16 479.38 479 [12] Quintozene (PCNB) 82-68-8 pentachloronitrobenzene<br>99.95 4233 41 417.47 417.58 417 [12] Silvex, methyl ester 4841-20-7 methyl 2-(2,4,5-trichlorophenoxy)propionate 99.40 7636 360.62 361.16 362 [14] 2,4,5-T methyl ester 1928-37-6 (2,4,5-trichlorophenoxy)acetic acid methyl ester 99.65 7281 361.89 363.35 363 [14] Tecnazene 117-18-O 1,2,4,5-tetrachloro-3-nitrobenzene<br>99.87 4650 373.22 373.22 373.55 372 [12] Tetradifon 116-29-O 1,2,4-trichloro-5-[(4-chlorophenyl)sulfonyl]benzene 99.93 6917 419.82 420.09 421 [12] 1,2,3-Trichlorobenzene 87-61-6 99.48 4124 322.91 323.75 326 [15] 1,3,5-Trichlorobenzene 108-70-3 99.81 4110 333.87 334.45 336 [15] ORGANOPHOSPHORUS COMPOUNDS Acephate 30560-19-1 O, S-dimethylacetylphosphoramidothioate<br>98.86 4928 363.92 30 98.86 4928 363.92 365.10 366 [12] Azinphos ethyl 2642-71-9 S-(3,4-dihydro-4-oxobenzo[d]-[1,2,3]-triazin-3-ylmethyl) O,O-diethylphosphorodithioate 98.31 6027 322.16 324.44 326 [12]











**99.68 10112 584.26 584.55 588 [12] Terbacil 4902-51-2 5-chloro-3-(l,l-dimethylethyl)-6-methyl-2,4(1H,3H)-pyrimidinedione 99.78 2989 448.04 449.48 450 [12]** 

## **UREAS**













The scan rate determination was a compromise between accuracy and analysis time [lo]. If the sample could be resolidified, it was remelted and the remelt results reported. Results for 273 compounds are summarized in Table 1. An example endotherm is shown in Fig. 1, and a van't Hoff plot in Fig. 2.

#### **RESULTS AND DISCUSSION**

Experimentally determined temperature onset (Exp.  $T$ ), heat of fusion  $(\Delta H_t)$ , and purity (% purity) determinations by DSC are presented in Table 1. Average values shown were calculated from two or more determinations. Theoretical temperature onset (Theo. T) for 100% pure compound, and literature melting points (Lit. m-p.) are included for comparison and usually agreed within 3 K. Compounds are grouped according to class and ordered alphabetically by common name. Many of the compounds are multifunctional, so the grouping by class cannot be completely unambiguous. The common name may be a trivial chemical name, a familiar trade name, or an abbreviation. Because most of these compounds are structurally complex, Chemical Abstracts Service (CAS) and International Union of Pure and Applied Chemistry (IUPAC) names are seldom used except for verification. The CAS names and CAS registration numbers are provided in Table 1 for clarity and structural verification.



**Fig. 1. DSC thermogram of fentin acetate.** 



**Fig. 2. van't Hoff plot of fentin acetate.** 

Temperature onset values were obtained from thermograms by taking the slope of the melting curve at the inflection point and extrapolating to the baseline. Most compounds reported in this study show a temperature onset very close to the literature melting point. However, a few compounds were observed to exhibit melting point depression. Possible reasons for the discrepancies may be related to the differences in the purity levels of compounds analyzed (see Fig. 3 and Table 2), or to the accuracy of the method used to determine literature values. It has long been recognized that compounds prone to sublimation should be analyzed in a sealed tube; such was the case in literature references for camphor and chloranil. Traditional literature melting points or ranges have been measured by observing the temperature at which melting begins and the temperature range over which the sample melts. Weaknesses in the traditional open and closed capillary tube methods include the possibility of inaccurate thermometers, melting point variation with ambient pressure, and hygroscopic variability,

Other identification techniques, such as mass spectrometry (MS), infrared (IR), or nuclear magnetic resonance (NMR) spectrometry, and/or elemental analysis were also used routinely for these compounds. Generally, two complementary techniques were applied, one such as DSC primarily to ascertain purity, and the other, such as MS, primarily to verify identity.

Some compounds exhibited anomalous behavior such as multiple peaks and unusual curve shapes resulting in failure of the DSC technique for certification of purity. Some considerations which may cause anomalous



**Fig. 3. Effect of added impurity on DSC thermogram.** 



National Institute of Standards and Technology, Standard Reference Material 1514 thermal analysis purity set. p-Aminobenzoic acid impurity in phenacetin

behavior are: (1) the presence of impurities which are immiscible in the liquid phase and melt at their own characteristic melting point; (2) solidsolid phase transitions such as a change in crystal structure; (3) a crystal acquiring rotational freedom at a lower temperature than that required for translational freedom [7,8,11]; (4) solids having significant vapor pressure; and (5) sublimation such as is observed for camphor when melted in an open capillary tube.

The presence of impurities can affect the thermogram in several ways. Multiple peaks and distorted curves are the most common indication of impurities. The assumption that an impurity is insoluble in the solid phase is not valid if the molecule can fit into the crystal lattice of the solid phase without causing distortion. This situation occurs in those cases where an impurity is similar in size to the principal component. For example, when hexachlorobenzene is added to pentachloronitrobenzene, the purity determination by DSC is falsely high [7]. Performing a quantitative analysis by another method helps to insure that the purity reported is a true value.

When compounds have more than one crystalline form, they can first be heated above the melting point of the lower melting form in an attempt to convert the lower melting crystals to the higher melting crystal form before a purity determination is performed. Bentazon, chlorfenson, and monuron are three compounds reported in this study which exhibit dimorphic behavior but could be successfully analyzed for purity by DSC.

Compounds having a low entropy of fusion ( $\Delta S_f < 5$  cal mol<sup>-1</sup> K<sup>-1</sup>) exhibit rotational motion in the solid form and need only to obtain translational motion on melting [8]. The  $\Delta S_f$  value can be calculated from DSC data ( $\Delta S_f = \Delta H_f / T_0$ ). Such compounds may readily form solid solutions

TABLE 2







**Fig. 5. DSC thermogram of dacthal acid.** 



**Fig. 6. DSC thermogram of duraset.** 

with impurities. Purity values determined for compounds with low entropies of fusion are likely to be suspect (e.g. results for ametryn are not included in Table 1). Additionally, acenaphthylene, 2-aminobutane hydrochloride, chromium trioxide, endosulfan I, fentin hydroxide, pentachlorophenol, 1,3-propane sulfone, terbacil and thiourea all had  $\Delta S_f < 8$  cal mol<sup>-1</sup> K<sup>-1</sup> and thus may have the potential of forming solid solutions with impurities, although they were included in Table 1. Experimental purity values could be higher than the true purity for these compounds.

Most compounds exhibited normal behavior. However, certain chemical classes exhibited noteworthy characteristics under DSC. Eight classes are discussed below.

#### *(1) Amides*

Compounds of this class usually can be analyzed by DSC if no other functional groups are present in the molecule, such as phenolic hydroxyl or carboxyl, which could react with the amide.

### (2) *Carbamates and ureas*

Compounds decomposing at or near their melting point are unsuitable for DSC. N-methylcarbamates, ureas, and thioureas tend to decompose readily.

Thiophanate melts, but has a decomposition temperature just above its melting point. Only one N-methylcarbamate tested, 2-chlorophenyl-N-methylcarbamate (CPMC), provided an acceptable thermogram. Another Nmethylcarbamate, 2-(sec-butyl)phenyl-N-methylcarbamate (BPMC), showed two small but distinct endotherms near its melting point (Fig. 4) and was not included in Table 1. The small amount of energy absorbed indicates the compound has already gained rotational energy, while the two peaks in the graph are probably caused either by the BPMC being polymorphic or by two separate compounds being present.

### (3) *Carboxylic acids*

Thirteen compounds of this class were analyzed successfully, having melting points ranging from 338 K to 475 K. If the melting point of the acid is higher, decomposition begins to occur. Higher-melting acids, such as dacthal diacid, which melts at about 605 K, often showed some decomposition, as evidenced by a sharp exotherm followed immediately by a complex and nonreproducible endotherm (Fig. 5). They were not included in Table 1.

Duraset, which contains ortho-carboxylic acid and amide groups, displayed evidence of reactivity. After melting at 422 K, a clear exotherm was seen at 428 K, possibly due to the functional groups reacting with one another (Fig. 6). Naptalam sodium salt also has an amide group ortho to the carboxylic acid salt and showed an exotherm immediately before the endotherm. These compounds were not included in Table 1.

### *(4) Organophosphorus compounds*

As was observed by Plato [8], organophosphorus compounds are difficult to resolidify under DSC conditions. Thus, results reported for many of these compounds are from initial melts only. However, a small amount of decomposition at the melting point may not be apparent until a second melt is performed. Some samples of dialifor, a phosphorodithioate, partially decomposed during the initial melt. The decomposition appeared on the remelt thermogram in the form of a distinct shoulder on the leading edge of the melting peak. The decomposition was not readily visible on the initial thermogram.

### *(5) Phenols*

Most of the compounds of this class were successfully analyzed by DSC. Failures occurred when other functional groups were present and could react with phenolic groups. Phenols with unusual thermograms included niclosamide with an amide group *ortho* to the phenolic group, and 3,6-dichlorosalicylic acid with a carboxylic acid group *ortho* to the phenolic group. The niclosamide thermogram (Fig. 7) showed a small heat of fusion with an abrupt leading edge in the melting peak and a long trailing edge. These two compounds were not included in Table 1.

### (6) *Triazines*

Some pesticide samples of the 1,3,5-triazine class show an unusual characteristic thermogram shape (see Fig. 8, atrazine). Presumably, ametryn was non-ideal because the experimentally determined entropy of fusion was less than 5 cal mol<sup>-1</sup> K<sup>-1</sup>. The leading edge of the melting curve showed an irregularity containing two or three bumps. Remelting the sample greatly reduced or eliminated this behavior. If the sample had been rapidly crystallized from solution, the s-triazine crystals might have internal mechanical (crystal lattice) stress, which appeared in the initial melting curve as the bumps noted above.

## (7) *Ethers*

Several ethers were determined successfully, such as methoxychlor, Silvex, 2-methyl-2-chlorophenoxyacetic acid (MCPA), 2-methyl-4-chlorophenoxy-



**Fig. 7. DSC thermogram of niclosamide.** 



Fig. 8. DSC thermogram of atrazine.



Fig. 9. DSC thermogram af PCP methyl ether.





Fig. 10. DSC thermogram of strychnine.

butyric acid (MCPB), 2-methyl-4-chlorophenoxypropionic acid (MCPP), 2,4-D and 2,4,5-T acids. The pentachlorophenol (PCP)-methyl ether (pentachloroanisole) thermogram showed two distinct endotherms at 377 K and 381 K (Fig. 9). Remelting the sample caused the peaks to sharpen. Analysis of this compound by GC/MS indicated very high purity. The PCP-methyl ether sample may have retained a dimorphic form. This compound was not included in Table 1.

### *(8) Alkaloids*

Many organic amines or alkaloids proved amenable to purity determination by DSC. However, strychnine produced a symmetrical peak with a convex baseline (Fig. 10), indicating m.p. 541-563 K (literature m.p. 544-563 K dependent on rate of heating). While this thermal behavior, lacking a normal endotherm and melting point, rendered our sample of strychnine not amenable to purity determination by DSC, a thermogram consistent with reported behavior was produced.

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